



Maratha Vidya Prasarak Samaj's

COMMERCE, MANAGEMENT AND COMPUTER SCIENCE (CMCS) COLLEGE, NASHIK

Udoji Maratha Boarding Campus, Gangapur Road, Nashik-422013

NAAC Accredited "B" Grade (CGPA 2.29)

B.Sc. Chemistry COURSE OUTCOMES

Subject	Course Outcomes (F.Y.B.Sc)
Physical Chemistry	<p>CO1 Students will be able to apply thermodynamic principles to physical and chemical process.</p> <p>CO2 Calculations of enthalpy, Bond energy, Bond dissociation energy, resonance energy.</p> <p>CO3 Variation of enthalpy with temperature –Kirchoff's equation.</p> <p>CO4 Third law of thermodynamic and its applications.</p> <p>CO5 Relation between Free energy and equilibrium and factors affecting on equilibrium constant.</p> <p>CO6 Gas equilibrium, equilibrium constant and molecular interpretation of equilibrium constant.</p> <p>CO7 Concept to ionization process occurred in acids, bases and pH scale.</p> <p>CO8 Related concepts such as Common ion effect hydrolysis constant, ionic product, solubility product.</p>
Organic Chemistry	<p>CO1 The students are expected to understand the fundamentals, principles, and recent developments in the subject area.</p> <p>CO2 It is expected to inspire and boost interest of the students towards chemistry as the main subject.</p> <p>CO3 To familiarize with current and recent developments in Chemistry.</p> <p>CO4 To create foundation for research and development in Chemistry.</p>

<p>Inorganic Chemistry</p>	<p>CO1 Various theories and principles applied to reveal atomic structure.</p> <p>CO2 Origin of quantum mechanics and its need to understand structure of hydrogen atom.</p> <p>CO3 Application of non-bonded lone pairs in shape of molecule.</p> <p>CO4 Explain rules for filling electrons in various orbitals- Aufbau's principle, Pauli exclusion principle, Hund's rule of maximum multiplicity.</p>
<p>Analytical Chemistry</p>	<p>CO1 Calculations of mole, molar concentrations and various units of concentrations which will be helpful for preparation of solution.</p> <p>CO2 Relation between molecular formula and empirical formula.</p> <p>CO3 Basics of chromatography and types of chromatography.</p> <p>CO4 Compare qualitative and quantitative analyses.</p>
<p>Chemistry Practical 1&2</p>	<p>CO1 Determination of physical constant: Melting point, Boiling point.</p> <p>CO2 drawing organic molecule and arrow pushing concepts.</p> <p>CO3 Strength of Acid and Base.</p> <p>CO4 Common names and IUPAC nomenclature system of chemicals.</p> <p>CO5 Name of Alkane, Alkanes, Alkenes and Alkynes.</p> <p>CO6 Preparation methods of Alkane, Alkenes and Alkynes including Hydrocarbons.</p> <p>CO7 Application of Huckel's rule of organic compounds to find the compounds are aromatic/ non aromatic.</p>



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